Infrared studies of species adsorbed in zeolites

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INFRARED STUDIES OF SPECIES ADSORBED IN ZEOLITES

BY

ASIA NASSER AL NOAIMI
(B.Sc. Qatar University)

A thesis submitted for the degree of Master of Science in the Department of Chemistry University of Durham

1981
Dedicated to my father Nasser Al Noaimi
Declaration

The work described in this thesis was carried out at Durham University and has not been submitted either wholly, or in part, at this or at any other university.
Acknowledgements

I would like to express my thanks to Dr. J. Howard for his advice so freely given throughout the period of research.

The help of various members of the technical staff is greatly appreciated. May I also thank my friends for their encouragement, and Mrs. R. Hart for typing this thesis.

Finally I would like to express my thanks to the Government of Qatar and Qatar University for their financial support.
The work described in this thesis is concerned with the investigation of the infrared study of zeolites and species adsorbed on zeolites.

In the first three chapters there are accounts on the zeolites structure and chemistry as well as the instrumentation and techniques which have been used.

In the fourth chapter a study of the formation of coke on CaY zeolite and the various factors which influence its presence on the zeolite is reported.

Chapter five is concerned with the presence of the hydronium ion formed by adsorbing water vapour on zeolites. Mn 4A zeolite can form the hydronium ion, while Co-4A does not.

Finally, in chapter six, we report a study of cyclopropane gas adsorbed on Mn(II) and Co(II) partially exchanged A zeolites.
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2. Experimental

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**Chapter 7: Conclusions and suggestions for future work**
1. **Zeolites**

Although zeolites are chemically more complex than some other adsorbents they do possess the advantages of high surface area and adsorption sites which are usually well defined. Furthermore, the strength of the zeolite-to-adsorbed molecule interaction can be altered in a variety of ways including ion-exchange and modification of the Si to Al ratio (1). It is possible, therefore, to study the interaction of a given adsorbate with a series of different cations, each held within the same framework as well as with the same cation in a range of frameworks.

Zeolites have strong affinity for polar or polarizable molecules and this property, combined with the internal adsorption characteristics, allows for purifications and separations to be performed using zeolites. In addition, the properties of zeolites can be altered still further by ion-exchange to provide a nearly limitless variety of products and potential uses (2-4).

Zeolites can also form novel chemical complexes like C$_3$H$_6$ (cyclopropane) or N$_3$H$_3$ bonded to transition metal ions (5). Note that N$_3$H$_3$ does not exist in the pure form and has not been observed outside a zeolite cage (6).

Numerous attempts are currently being made to use zeolites to solve, or improve the solution of technological problems. New types of zeolites and variations on existing types are being synthesized at present and many more will undoubtedly be discovered hopefully in the future, scientists
will be able to synthesize a zeolite-type material to meet particular applications.

The first applications of zeolites in industry was in separations, but this was followed by important catalytic uses. Recently, they have also been applied in solving environmental problems (3-9).

2. **Spectroscopic studies**

Spectroscopic techniques can ideally give information about the nature of surfaces and species adsorbed on surfaces. It has generally proved difficult to obtain the Raman spectra of zeolites, and particularly of zeolites containing adsorbed species. Zeolites themselves give rise to very weak Raman scattering and their fluorescence background is often high (9); this fluorescence is due to the presence of transition metal ions or of trace quantities of organic impurities (10). Similarly, great difficulty has been experienced with far-infrared studies (11). Apart from the intrinsic experimental difficulties, the low frequency normal modes of the adsorbed species appear to yield bands which are of very low intensity relative to the vibrations of the framework.

Infrared spectroscopy can give definite information on the structure and surface properties of zeolites and of how these are modified by various treatments. Changes in the spectra of the zeolite itself and of molecules adsorbed on the surface can yield direct information about the surface, the position of adsorbed molecules and the interaction between them (11-18). Infrared spectroscopy is a most sensitive structural technique (3).
The detailed absorption bands of the solid support and of the adsorbed species are not always all observed. Sometimes the absorption bands of prime interest are obscured by strong bands associated with vibrations of the zeolite.

Despite the problems that may arise, infrared spectroscopy remains a powerful method of studying adsorbed species.
References


CHAPTER 2
THE STRUCTURE AND CHEMISTRY
OF ZEOLITES

1. Zeolites structures and properties
   a. General structure

   Zeolites are framework aluminosilicates based on a three-dimensional network of $\text{AlO}_4$ and $\text{SiO}_4$ tetrahedra linked to each other by corner sharing of oxygens. Since the Al atoms are trivalent, each $\text{AlO}_4$ unit is negatively charged and the charge on these units is balanced by framework cations. In the hydrated zeolite, the cations are usually quite mobile in the channels and therefore can be exchanged, at least to some extent, by other cations (1-9).

   Zeolites may be represented by the empirical formula:

   $$\frac{M_{2/n}}{O} \cdot \text{Al}_2\text{O}_3 \cdot X\text{SiO}_2 \cdot Y\text{H}_2\text{O} (2,6)$$

   Where $n$ is the valence of the cation $M$ and $X \geq 2$ since $\text{AlO}_4$ tetrahedra are assumed to be joined to $\text{SiO}_4$ tetrahedra. A large number of both natural and synthetic zeolites are known. The structures of these belong to distinct groups with common building blocks; zeolites A, X and Y have frameworks consisting of linked truncated octahedra called sodalite units (Fig.1a) which are characteristic of the structure of sodalite. These are composed of 24(Al, Si) ions interconnected with 36 oxygen anions, and contain eight hexagonal and six square faces.

   In zeolites X and Y, the framework consists of a tetrahedral arrangement of sodalite units linked by six
bridging oxygen atoms. The structure is very open with 12Å diameter almost spherical cavities (supercages) (Fig. 1b), each of which is surrounded tetrahedrally by four similar cavities separated one from another by constrictions 8-9Å in diameter, formed by distorted "12-rings" of oxygen atoms. The unit cell composition of zeolite X is Na₈₆·(AlO₂)₈₆·(SiO₂)₁₀₀·26₄ H₂O.

The basic framework for zeolite Y is the same as that of zeolite X, but the Si/Al ratio is higher, generally ranging from 1.5 to 3.0.

In zeolite A the truncated octahedron is linked to its neighbour by four bridging oxygen ions (10,11); this configuration gives a roughly spherical internal cavity called the \( \alpha \)-cage (Fig.1c) 11.4Å in diameter. This may be entered through six approximately circular windows 4.2Å in diameter, the circumferences of which contain eight oxygen atoms and are hence known as "8-rings". A second set of voids (\( \beta \)-cages:6.6Å in diameter) is contained within the sodalite unit themselves. These interconnect with the \( \alpha \)-cages through a distorted "6-ring" of oxygen atoms 2.2Å in diameter. There may, therefore, be considered to be two interconnecting pore systems, one of diameter 11.4Å with 4.2Å constrictions, and the other of alternate 11.4Å and 6.6Å cavities separated by 2.2Å constrictions.

b. Cation positions

Cations are usually distributed over a range of different sites within a particular zeolite framework (2). According to X-ray structural analysis, the distribution
SODALITE

(a)

ZEOLITE 13X

(b)

ZEOLITE A

(c)

Fig. 1
of cations is altered after the adsorption of water or treatment at different temperatures (13). In general the cations distribute themselves amongst the sites so as to minimise the free-energy and there is usually only partial occupancy of many of the available sites. This distribution is a function of temperature, the cationic species, the state of hydration and the zeolite Si/Al ratio (1).

In the mid-1960s three main types of site were generally recognized, later the number of defined cation sites increased and these are described as follows (Fig. 2).

- Site I situated in the centre of the hexagonal prism.
- Site I' on a triad axis displaced into the sodalite cage from the hexagonal face shared by the sodalite cage and the hexagonal prism.
- Site II' on a triad axis displaced into the sodalite cage from the open six membered ring of the sodalite unit.
- Site II displaced from the open six membered ring into the supercage.
- Site III displaced into the supercage from bridging four membered rings and
- Site IV very nearly at the centre of the twelve membered rings separating the supercages (1-4, 14).

It is reasonable, therefore, that those positions most favourable to the cation under consideration will fill first, and subsequently any remaining cations will occupy sites in order of increasing unsuitability.
Fig. 3

The cation sites in zeolite Y

The lines represent the bonds between the T atoms (Si, Al) and the oxygen atoms. The oxygen atoms are located between the T atoms and are not necessarily on the line joining them.
The possible cation positions within type Y and A zeolite are shown in Tables 1 and 2 respectively.

Table 1: The cation positions in dehydrated NaY zeolites (2)

<table>
<thead>
<tr>
<th>Types of Sites</th>
<th>Max. available</th>
<th>Site occupancy</th>
</tr>
</thead>
<tbody>
<tr>
<td>I</td>
<td>16</td>
<td>7.5</td>
</tr>
<tr>
<td>I'</td>
<td>32</td>
<td>19.5</td>
</tr>
<tr>
<td>II'</td>
<td>32</td>
<td>-</td>
</tr>
<tr>
<td>II</td>
<td>32</td>
<td>30</td>
</tr>
<tr>
<td>U</td>
<td>8</td>
<td>-</td>
</tr>
</tbody>
</table>

c. Ion exchange

Ion exchange in zeolites has been widely studied. Cation selectivities in zeolites, however, do not follow the typical rules (2) that are evidenced by other inorganic and organic ion exchangers. Zeolite structures have unique features which lead to unusual types of cation selectivity and sieving. The cation exchange behaviour of zeolites depends upon the (2,3):

a) Nature of the cation species, the cation size, both anhydrous and hydrated, and cation charge,
b) Temperature,
c) Concentration of the cation species in solution,
d) The anion species associated with the cation in solution,
e) Structure and characteristics of the particular zeolite.
Table 2: Cation Locations in Type NaA zeolite

<table>
<thead>
<tr>
<th>Position</th>
<th>Designation</th>
<th>Number</th>
</tr>
</thead>
<tbody>
<tr>
<td>In an 8-ring</td>
<td>S1</td>
<td>3 per unit cell</td>
</tr>
<tr>
<td>Adjacent to an 8-ring, but displaced into a 26-hedron (α-cage).</td>
<td>S1*</td>
<td></td>
</tr>
<tr>
<td>In a 6-ring</td>
<td>S2</td>
<td>8 per unit cell</td>
</tr>
<tr>
<td>Adjacent to a 6-ring, but displaced into a 26-hedron (α-cage).</td>
<td>S2*</td>
<td></td>
</tr>
<tr>
<td>Adjacent to a 6-ring, but displaced into a sodalite β-cage.</td>
<td>S2'</td>
<td></td>
</tr>
<tr>
<td>Against a 4-ring forming one of the ribs of a 26-hedron (α-cage).</td>
<td>S3</td>
<td>12 per unit cell</td>
</tr>
<tr>
<td>In the centre of a sodalite β-cage</td>
<td>SU</td>
<td>1 per unit cell</td>
</tr>
<tr>
<td>In the centre of a 26-hedron (α-cage).</td>
<td>S4</td>
<td>1 per unit cell</td>
</tr>
</tbody>
</table>

Generally it is easier for a cation to diffuse down a large channel and the temperature at which the exchange is carried out is often of great importance. The ion exchange capacity of zeolites is large and can approach the capacity of the best resin ion exchangers. This ability to undergo reversible cation exchange is one of the most important properties of zeolites, the most striking modification that can be made in this respect is to change the molecular sieve properties. For example, the replacement of Na\(^{+}\) ions in Linde A by K\(^{+}\) ions cause a decrease in the \(O_2\) sorption to essentially zero (17). Also, changing the cation for
one of higher charge in a zeolite may effectively enlarge the pore openings by diminishing the cation population and by resiting the cations which are normally located near these openings. In zeolite A, divalent ion exchange opens the aperture to full diameter whereas exchange with a larger univalent ion diminishes the aperture size. Potassium ion exchange in zeolite A reduces the effective adsorption pore size to the point where only small polar molecules are adsorbed (2).

d. Adsorption

In contrast with other molecular sieves, zeolites have pores of uniform size which are uniquely determined by the structure of the crystal. Fig. 3 shows for comparison the pore size distribution for zeolites 4A and 13X, silica gel and activated carbon (10,11,13,14,16-19). The dehydrated crystalline zeolites are the most important molecular sieves (6). These materials have a high internal surface area available for adsorption due to the channels or pores which uniformly penetrate the entire volume of the solid. The external surface of the adsorbent particles contributes only a small amount of the total available surface area. The first experimental observations of the adsorption of gases on zeolites and their behaviour as molecular sieves were conducted on the zeolite minerals.

The adsorbed phase disperses through the internal voids of the crystal without displacing any atoms which make up the permanent crystal structure. There are several factors which influence the rate of adsorption in zeolites such as pressure, temperature, particle size and the conditions of
Fig. 3  pore size distribution of various molecular sieves

a. type 4A zeolite
b. type 13X zeolite
c. silica gel
d. activated carbon
dehydration. The adsorption of a molecule depends partially upon its polarity and polarizability together with variations in the chemical composition of the zeolite (6).

The main channels by which diffusion occurs in zeolites are formed by cavities connected by apertures, the walls of the channels, which form the internal surface, contain a regular array of cations that bear a charge depending upon the inadequacy of the local coordination or screening by the oxygen ions of the framework. Upon dehydration and removal of water from the channels, the remaining cations are stranded at channel intersections or on channel walls where they inhibit diffusion of other molecules.

e. Catalytic activity

Zeolites offer unique opportunities for studies of heterogeneous catalysis because of their catalytic, structural and ion-exchange properties. These crystalline aluminosilicates, in appropriate ion exchanged forms, can offer sizeable activity enhancements and selectivity alterations for certain reactions (1) when compared with amorphous silica-alumina, although not necessarily for all reactions (20).

The existence of several different zeolite framework structures, each of which can be rationally modified by means of controlled ion-exchange, offer a range of catalytic properties. There are several factors which effect the catalytic activity including (a) the number of active sites, (b) the size and shape of adsorbed molecules, and (c) the rate of diffusion (1). Various guest molecules can also
significantly change the catalytic properties of the zeolites. It is known that the addition of small amounts of water increases the activity of zeolites for carbonium-ion type reactions, cracking alkylation and isomerization (1). The addition of water molecules to X and Y zeolites with monovalent ions increases the isomerization of cyclopropane (21).

2. Applications

The growth of the zeolite industry has been extremely slow considering that it is 25 years since commercially mineable deposits were located (22).

The first application of dehydrated zeolites as molecular sieves was in the separation of gas mixtures. The commercial uses of zeolites in adsorption are summarized in Table 3.

Table 3: Adsorption applications (23)

Regenerative

separations based on sieving

separations based on selectivity

purification - bulk separations.

Non-regenerative

Drying refrigerants.

cryosorption.

One of the earliest applications in ion exchange was in the removal and purification of caesium and strontium radio-isotopes. Uses of zeolites in ion exchange are listed in Table 4 (23).
Table 4: Ion-exchange application

Removal of NH$_4^+$ from waste water
Metal separation
Radio-isotope removal and storage
Detergent builder
Artificial kidney dialysate regeneration
Aquaculture - NH$_4^+$ removal
Ruminant feeding on non-protein nitrogen
Ion-exchange fertilizers.

More zeolites are currently used in catalysis than in any other application, e.g. as cracking catalysts for converting petroleum into lighter fractions including gasoline and fuel oil. Petrochemical feeds have provided the major market for these materials.

Some applications of zeolites in catalysis are given in Table 5.

Table 5: Catalysis applications

Hydrocarbon conversion
Alkylation
Cracking
Hydrocracking
Isomerization
Hydrogenation and dehydrogenation
Hydro-dealkylation
Methanation
Shape-selective reforming
Dehydration
Organic Catalysis
Inorganic reactions

$H_2S$ oxidation

3. Application of infrared spectroscopy to the study of zeolites

Infrared spectroscopy has been of the greatest value in the determination of the structure of molecules. It has been very widely used in research in the field of surface chemistry and in the study of adsorbed species.

a. Infrared studies of zeolite frameworks

Infrared spectroscopy has been extensively used to elucidate the framework structure of zeolites. For instance, Zhdanov et al (24) have applied infrared spectroscopy to the study of the frameworks of a series of X and Y zeolites as a function of Si/Al ratio, cation type and state of hydration. In the region investigated, 400-800 cm$^{-1}$, they identified two types of vibration: the first due to internal vibrations of the $(Al,Si)_4O_{12}$ tetrahedron, which is the primary unit of the structure and which is not sensitive to other structural variation, and the second which may be related to the linkages between tetrahedra. The most detailed study of the framework structures of zeolites and their classification to groups has been made by Flanigen et al (25). The study was made in the 1300-200 cm$^{-1}$ region since this range contains the fundamental vibrations of the $(Si,Al)O_4$ tetrahedra.

A summary of some infrared assignments of framework vibrations are given in Table 6. This is data for NaY (Fig.5) zeolite and is intended to illustrate the assignment scheme (25). In the spectra shown, the internal
tetrahedral linkages are drawn in full line, and the structure sensitive external vibrations with broken lines \((8,10,26,27,28)\).

Table 6: Infrared assignments of zeolite framework vibrations \((2.5)\)

<table>
<thead>
<tr>
<th>Internal tetrahedra</th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>Asym. stretch</td>
<td>1250 - 950</td>
</tr>
<tr>
<td>Sym. stretch</td>
<td>720 - 650</td>
</tr>
<tr>
<td>T-O bend</td>
<td>500 - 420</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>External linkages</th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>Double ring</td>
<td>650 - 500</td>
</tr>
<tr>
<td>Pore opening</td>
<td>420 - 300</td>
</tr>
<tr>
<td>Sym. stretch</td>
<td>750 - 820</td>
</tr>
<tr>
<td>Asym. stretch</td>
<td>1150 - 1050 sh</td>
</tr>
</tbody>
</table>

Fig. (5) Infrared assignments illustrated with the spectrum of zeolite Y, Si/Al of 2.5 \((25)\).

The full line indicates vibrations due to the internal tetrahedral linkages and the broken line the structure sensitive external vibrations.
The hydroxyl groups in zeolites

The hydroxyl groups have been extensively studied by infrared spectroscopy. Three types of hydroxyl groups were detected by Bertsch and Habgood (39-32). These measurements provide information concerning the location of the hydroxyl groups, their functionality and (28,29,30) information concerning the location of the cations themselves. The bands at 3660 and 3550 cm$^{-1}$ are due to the most important hydroxyl groups and are referred to as the high frequency and low frequency bands, respectively.

Infrared provide useful information as to how these surfaces are modified by various treatments.

b. Adsorption of molecules

A large number of studies and reviews have been published of molecules adsorbed on zeolites (5,8,16,32-34). The earliest study was of benzene (33) over Na and CaX zeolites by Kiselev et al. The C-C stretching vibration band at 1486 cm$^{-1}$ was very much stronger than the CH stretching band. A detailed study of the adsorption of ethylene has been reported (34). Very large intensity changes, larger than those previously observed with benzene (33) were found. Eberly studied the interaction between olefins and zeolites at high temperatures. The behaviour of adsorbed species at high temperatures is particularly useful in relation to catalysis, which is often conducted under these conditions (35-37). Other applications of infrared spectroscopy to the study of zeolites include:

1) The $\text{SiO}_2/\text{Al}_2\text{O}_3$ ratio

The frequency shift of the infrared stretching bands with the variation in Si/Al ratio in the tetrahedral frameworks has been reported (38,39). From infrared spectra of the synthetic sodium zeolites A, X, Y erionite and mordenite, Shikumove et al, drew a quantitative relationship
between the position of the main asymmetric stretch band and zeolite Si/Al ratio, Flanigen and Grose (40 used mid-infrared spectroscopy to characterize the framework composition of phosphate zeolites and establish proof of phosphorous substitution in the framework. The observation and assignment of Infrared frequencies can generally be applied to prove that framework substitution has occurred.

(2) Cation type and cation sites

The nature and crystallographic sites of cations in zeolites have been reported to be reflected in the infrared spectra. Zhdanov et al (24) showed the strong sensitivity of framework infrared vibrations to cation type and charge for a series of Na, Sr and Ca ion-exchanged forms of zeolite X. Flanigen et al (25) noted that dehydration of the synthetic forms of zeolites A, X, Y, L and Α, all containing alkali metal cations, showed only minor spectral changes; however, they reported significant changes in spectra of dehydrated zeolites in polyvalent cation forms, such as CaY, as shown in Fig.6. The spectral changes were interpreted as due to cation movement or resiting of cations as a result of dehydration, dehydroxylation and rehydration reactions, and it shows the migration of Ca$^{2+}$ cations from inside of the sodalite unit into a position near the centre of the D-6 ring (site I).

Brodskii, Zhdanov and Starevich (41,42) studied the effect of dehydration on a series of alkali-metal cation exchanged forms of zeolite X and Y with varying Si/Al ratios. Their spectral data shows that spectral changes are sensitive to the nature of the substituting cation and
Fig. 6 Infrared spectra of CaY zeolite (Si/Al of 2.5) after dehydration, dehydroxylation and rehydration.
the cation site. They also showed that the framework frequency is proportional to $m^{-2/3} R^{3/2}$ where $m$ is cation mass and $R$ is cation radius (42).

(3) **Zeolite synthesis mechanism**

Infrared spectroscopy can also apply to investigations of the mechanism of crystallization of zeolites from hydrous aluminosilicate gels or other aluminosilicate systems, such as the crystallisation sequence for NaY zeolite from a sodium aluminosilicate gel (2).

(4) **Catalytic system**

The application of infrared spectroscopy to the catalytic aspects of zeolites has been in two major areas.

a. the elucidation of the structural groups in zeolites and their properties as they pertain to catalytic centres, and

b. the observation of catalytic reactions on zeolites while the reactions are actually occurring on the zeolite surface. The surface of decationated Y zeolite was studied during cumene cracking on the zeolite (43), and it was found that the 3550 cm$^{-1}$ band is not influenced by the cumene while the 3640 cm$^{-1}$ band decreased in intensity with time. As the temperature was raised to over 325°C, however, the 3550 cm$^{-1}$ band intensity also decreased. This indicated that no interaction of the cumene with the species responsible for this band occurs below 325°C and that these sites are probably not active catalytic centres. The species producing the 3640 cm$^{-1}$ band are the active sites.

Although many studies have been reported on catalytic reactions over zeolites, the particular sites responsible
for reaction still remain uncertain in all but a few cases.
References


32. Yates, D.J.C., Catalysis Reviews 2 (1),113 (1968).


CHAPTER 3
TECHNIQUES AND INSTRUMENTATION

1. Infrared Spectrophotometer:

A Perkin-Elmer 580B infrared spectrophotometer was used. It is a double beam, ratio recording instrument, which provides a continuous record of the transmission of a sample as a function of frequency. (1) The source is a ceramic tube, it is heated to 1200°C and produces a continuous spectrum of electromagnetic radiation.

The radiation from the source is focused on a baffle by a toroid mirror, the baffle ensures that radiation from only a limited surface area of the source is admitted to the optical system thus minimizing sample heating. The baffle image is focused on to the first chopper mirror which rotates, dividing the source energy into sample and reference beams. The two beams are focused into the sample compartment. After the sample compartment the alternate pulses of radiation from the two beams are combined by the action of a second chopper, which has reflective first and third quadrants. During the first and second quadrants the second chopper receives energy via the sample and reference beams respectively. During the third and fourth quadrants chopper 1 cuts off the source energy so that any energy appearing at the second chopper is due to re-radiation effects from the sample compartment, then the two beams are combined into a single beam of alternating segments (1). When the beams are of equal intensity, the instrument is at an optical null. The recording pen is then at 100% T when no sample is present (2-4). The beam is then focused on the monochromator entrance slit. The slit restricts
the radiation passing through it to a narrow wavenumber band
the mean of which corresponds to the wavenumber at which the
measurement is being made. The effective bandwidth varies
with the wave number setting of the instrument since the slit
width is programmed to maintain approximately constant energy
at the detector over most of the wavenumber range.

The detector consists of a thermocouple within an
evacuated housing at the focus of an on-axis ellipsoid mirror,
the radiant energy leaving the exit slit is therefore
focused by the ellipsoid as an image of the slit reduced
in linear dimensions by a factor of eight. A caesium iodide
lens on the thermocouple assembly further reduces the
dimensions of the slit image falling on the target.

The alternating signal from the detector is amplified
and then demodulated by the signal processing electronics
to give separate sample and reference beam signals, which
are compensated for the effects of thermal re-radiation from
the sample compartment. The ratiometer produces the ratio
of the two signals, which corresponds to the transmittance
value of the sample. The ratio signal is then filtered to
reduce the noise level, and baseline adjustment, offsetting
and scaling operations are performed on the signal in the
ordinate functions unit. The signal is then supplied to the
recorder.

A brief outline of the operation of the PE 580 infrared
spectrophotometer is shown in the block diagram in Fig (1).

2. Infrared data station:

The Perkin-Elmer infrared data station is designed for
use with 580 B infrared spectrophotometers and other models.
Fig. 1 A block diagram of PE 580.
The data station is provided in three modules, (a) the visual display unit, which is used to display a spectra, (b) the keyboard, which is used to enter system commands, and (c) the data processing module which houses the system electronics and two micro-floppy disc drives.

When the data station is connected to the spectrophotometer an integrated infrared data system results, while independent use of the spectrophotometer is possible.

A block diagram of the system is shown in Fig (2) the system software provides the step-by-step instructions to the data station which in turn controls the spectrometer. The data processing module sorts the instructions to the instrument and accepts data from it while the working memory in the data processing module stores the spectral data for manipulation and modification by the applications programme routines. The data then may be saved on a bulk storage disc.

The PE 580 applications programme is provided on one of the microfloppy discs supplied with the system. This programme permits acquisition of spectra from the spectrophotometers, display of the spectra on the visual display unit, and subsequent storage of the spectra on disc.

There are twenty-four special function keys provided at the top of the keyboard module. One of the most important keys, which is often used, is the ABEX command. This command is equivalent to running a second spectrum with a sample of increased or decreased concentration. It is generally used by us to expand spectra of highly absorbing samples and the expansion factor may be entered as any value > 1. If the spectrum is stored in transmission (T%) each point is converted...
Fig. 2 system block diagram
to its corresponding absorbance value and the expansion is preformed on the spectrum and then reconverted to T%. If the spectrum has previously been converted to absorbance (A), it is adjusted to 0.0A and multiplied by the factor.

Reference (5) gives the detailed information about the function keys.

3. Thermogravometric analysis:

Thermogravometric analysis, (TGA) has been carried out using a TGA 750 to detect the changes of zeolite weight as a function of temperature. TGA provides the analyst with a quantitative measurement of any weight change associated with temperature. Changes in weight are a result of the rupture or formation of various physical and chemical bonds at elevated temperatures which lead to the evolution of volatile products. TGA curves are characteristic for a given compound because of the unique sequence of the reactions (6).

In this study, TGA has been used mainly to detect the loss of water molecules from zeolites as a function of temperature.

4. X-ray powder photography:

This method has been used to distinguish between the structure of X and Y zeolites in the powder form, and to test for sample decomposition. Table (1) contains a comparison of the literature data (7) on X and Y zeolite, compared with my values for the same zeolites. Cu K\text{α} was used as the source of radiation(7).

\[
K_{\text{α}} = \frac{2K_{\text{α1}} + K_{\text{α2}}}{3}
\]

For the copper target.

\[
CuK_{\text{α}} = 1.54178 \ \AA
\]
Table 1 - X-ray data ($d, R$) for X and Y zeolites

<table>
<thead>
<tr>
<th>NaX zeolite ref (7) $d(Å)$</th>
<th>NaY zeolite ref (7) $d(Å)$</th>
<th>NaX zeolite our data $d(Å)$</th>
<th>NaY zeolite our data $d(Å)$</th>
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<tr>
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<td>8.82</td>
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<td>2.86</td>
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</tbody>
</table>
5. **Infrared cell:**

A purpose built infrared cell was used for studying the infrared spectra of the adsorbed gases. A diagram of this cell is shown in Figs. 3a, 3b. The cell was made from stainless steel and was connected to a transformer which controls the temperature of the disc holder. The cell is also connected to outlets, for admission gases and for evacuations at (a). The head of the cell is a "Vacuum Generator" electrical lead through with 70mm UHV flange.

Full details of the dimensions of the cell are shown in Figs. 3a and 3b.

Gases are admitted to the cell using a valve in the vacuum system and the pressure of the gas is measured using a pirani gauge. For liquids a flask with two rotafloors were used to join between the cell and the valve of the vacuum line.

The cell can be heated in situ to 400°C, by using molybdenum wire, as the heating element. The spectra of the sample in this cell could therefore be recorded at any temperature between ambient and 400°C. A water jacket is shown at (b) through which water is circulated to prevent the windows from getting too hot.

KRS-5 (ThI, ThBr) windows were used, since these have high transmission in the near infrared region (20,000 - 250 cm\(^{-1}\)).

6. **Vacuum systems:**

The vacuum line used was metal made of stainless steel except for the gas handling part (which was of glass). Obvious requirements include a pirani and hot cathode ionization gauges. The pumps include the mechanical rotary
Fig. 3a side view of the cell

- bolt hole
- to electrical connection
- electrical lead through
- Mo wire
- thermocouple
- 52.5mm
- 42mm
- 50mm
- 60.5mm
- 74mm
- 76mm
- 53mm
- 48mm
- for O-ring
disc holder
knurled ring
water jacket
Fig. 3b  front view of the cell

- Mo wire
- water jacket
- disc holder

- Electrical connection
- Electrical lead through
oil pump which can give a vacuum down to $10^{-3}$ torr, and a turbo-molecular pump which goes to $10^{-7}$ torr or lower.

A block diagram which shows the whole system is shown in Fig.4

7. **Sample preparations:**

**Gases:**
The infrared spectra of the gases were obtained at a range of pressures in the metal cell without the zeolite disc.

**Liquids:**
These may be examined directly in very thin layers. KBr windows were used to obtain the spectra of the liquids.

**Solids:**
Transparent discs were made by compressing the zeolite with KBr or as a self-supporting disc. 10 mg of the zeolite were pressed in the die which has a diameter of 12 mm. After forming a self-supporting disc of the zeolite the sample was then inserted in the cell and left to evacuate until the pressure reaches $10^{-5}$ torr or less. The sample was then heated slowly until most of the water bands were removed. During heating the pressure increases initially but after the pressure decreased to $10^{-5}$ torr, the sample was cooled to the desired temperature, and the relevant adsorbate admitted.
Fig. 4  block diagram of the vacuum system
References


5. Instructions, Perkin-Elmer infrared data station PE 580 application programme, 1980. Perkin-Elmer Ltd


CHAPTER 4

INFRARED SPECTROSCOPIC INVESTIGATION

RELATING TO COKE FORMATION ON ZEOLITES:

ADSORPTION OF HEXENE-1 AND N-HEXANE

1. Introduction

Eisenbach and Gallei (1) studied the adsorption of hexene-1 and n-hexane on to Y type zeolites (calcium and ammonium exchanged) in order to elucidate the interaction of the functional hydroxyl groups on both the internal and external surfaces of the zeolite with these hydrocarbons.

By heating the CaY and NH₄Y zeolites with the hydrocarbons, they observed the formation of coke which lowered the activity of the catalyst due to the poisoning of the active catalytic centres. As the coke band increased in intensity, some of the hydroxyl bands decreased. As a result the authors correlated the active sites for coke formation with these hydroxyl groups. The identification of a particular site responsible for coking makes the work both interesting and industrially relevant and represent a new application of infrared spectroscopy.

Many studies of coke formation in zeolite e.g. kinetics, mechanism and the effect on the properties of catalyst have been made by a number of workers (2-9). Some of the work (10) was concerned with catalyst poisoning, which occurs when the hydrocarbon converts over the zeolite and gradually deposits coke on the surface.

None of these studies however elucidate the correlation
of coking activity with a particular site in the catalyst. In contrast with earlier work, this paper of Gallei and Eisenbach (1) links the interaction of the hydrocarbons with particular hydroxyl groups in the zeolite catalyst.

Coke formation involves condensation-hydrogen elimination reactions in which the H/C ratio of the coke and its degree of unsaturation decrease as the time spent by the catalyst particle in the reaction increases. Simultaneously with the deposit of coke on the catalyst the surface area decreased which indicates blocking of pores and active sites. This shortens the life of the catalyst. Deposits of coke at the pores lying on the outer surface of the catalyst prevent entry of reactants to the zeolites framework. In some cases regeneration is possible by blowing oxygen over the surface to reduce the coke.

In this chapter there will be a comparison between our results and those of Eisenbach and Gallei for hexene-1 and n-hexane adsorbed on calcium exchanged Y zeolite.

2. Experimental:

Materials

CaY and NH4Y zeolites were prepared from NaY zeolite (Union Carbide Corporation) by ion-exchange with 0.1 and 0.3N solutions of analytical grade CaCl2·2H2O, NH4NO3 (Analar) and Ca(NO3)2 (Hopkins and Williams). The ion-exchanges were carried out at ambient temperature, and the zeolites were analysed for Na,Al and either Ca or NH4⁺ as appropriate, using atomic absorption techniques.

Eisenbach and Gallei presented their results as percentages of the oxides (Al2O3, Na2O, CaO) and NH4⁺ and
so, to compare their results with ours, we have represented our data in the same way. Gallei et al. also assumed that their samples had zero water content. Since the zeolites were not dehydrated before analysis the percentage figures that they present cannot represent the true compositions of the zeolites. For more details see the analysis of the zeolites in Table 1. Analytical grade hexene-1 and n-hexane were used after distillation and storage over molecular sieve 4A.

**Procedures**

10mg of the CaY zeolite was pressed in the die (12mm diameter). As the disc was degassed by heating and pumping, spectra were obtained at a range of temperatures. Spectra of CaY zeolite were recorded at ambient temperature, 493 and 623K. The zeolite sample was then cooled to room temperature, the spectrum recorded, the hydrocarbon added and the spectrum obtained once more. The sample was then evacuated for 10 minutes, and the spectrum recorded. The zeolite was then exposed to the hydrocarbon once more at ambient temperature and the sample heated slowly to 473-573K and spectra recorded at a variety of temperatures.

3. **Results and discussions**

a. **Ion exchange**

As explained earlier, Eisenbach and Gallei assumed that their samples have zero water content but this is invalid because the zeolite was not dehydrated before analysis. In order to compare our data with theirs, we have calculated the ratios Al$_2$O$_3$/Na$_2$O, Al$_2$O$_3$/CaO and Al$_2$O$_3$/NH$_4^+$, since these are not affected by the presence of water.
<table>
<thead>
<tr>
<th>Sample number</th>
<th>Na\textsubscript{Y}(a)</th>
<th>102.0</th>
<th>62.0</th>
<th>-</th>
<th>-</th>
<th>-</th>
<th>1.65</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>Na\textsubscript{Y}</td>
<td>29.505</td>
<td>18.303</td>
<td>-</td>
<td>-</td>
<td>-</td>
<td>1.6</td>
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<tr>
<td>2</td>
<td>Ca\textsubscript{Y}</td>
<td>0.1</td>
<td>28.378</td>
<td>7.305</td>
<td>5.404</td>
<td>exact amount</td>
<td>25</td>
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<td>Ca\textsubscript{Y}</td>
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<td>29.429</td>
<td>6.946</td>
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<td>4</td>
<td>Ca\textsubscript{Y}</td>
<td>0.3</td>
<td>31.58</td>
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<td>6.72</td>
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<td>35</td>
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<tr>
<td>5</td>
<td>NH\textsubscript{4}\textsubscript{Y}</td>
<td>0.1</td>
<td>29.31</td>
<td>8.922</td>
<td>-</td>
<td>2.25</td>
<td>exact amount</td>
</tr>
<tr>
<td>6</td>
<td>NH\textsubscript{4}\textsubscript{Y}</td>
<td>0.1</td>
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<td>7.477</td>
<td>-</td>
<td>3.342</td>
<td>excess</td>
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<tr>
<td>7</td>
<td>NH\textsubscript{4}\textsubscript{Y}</td>
<td>0.3</td>
<td>29.84</td>
<td>6.613</td>
<td>-</td>
<td>3.6</td>
<td>excess</td>
</tr>
</tbody>
</table>

(a) Calculated from the standard formula, Na\textsubscript{56}(AlO\textsubscript{2})\textsubscript{56}(SiO\textsubscript{2})\textsubscript{136} \cdot 250 H\textsubscript{2}O

b. normality
Eisenbach and Gallei (1) prepared their zeolites from the Na form of faujasite by ion exchange with 0.1N solutions of NH$_4$NO$_3$ and Ca(NO$_3$)$_2$ and their analysis data is given in Table 2. By comparing Tables 1 and 2, it can be seen that the two sets of results are very different. It should be noted that the samples were prepared by the same methods, except that we do not have any information on how long Eisenbach and Gallei left their ion exchange to take place.

From Table 1, one can see that 0.1N solution gives a lower degree of ion exchange than that of Eisenbach and Gallei. In view of this, we carried out ion exchange in 0.3N solutions to try to reach the same degree of ion exchange that they reached. In both cases, however, the degree of exchange we obtained is less than that of Eisenbach and Gallei although they used only 0.1N solutions. In fact, the published analytical data of Eisenbach and Gallei is incorrect, and they have sent us a new table containing correct values (Table 3). Unfortunately this data does not agree with ours either.

Eisenbach et al (11) suggest that the zeolite which we used in our experiments is not Y zeolite but X. By obtaining the X-ray powder photograph of the sample, we have shown that the zeolite is indeed of type Y and not X (Table 4).

Several papers (13-16) show a range of values of the degree of ion exchange of CaY zeolite, when prepared by different methods. None of these results agree with ours or those of Eisenbach et al (1).
Table 2: Analysis of Zeolites by Eisenbach and Gallei (1)

<table>
<thead>
<tr>
<th>Sample</th>
<th>SiO$_2$ Wt %</th>
<th>Al$_2$O$_3$ Wt %</th>
<th>Na$_2$O Wt %</th>
<th>CaO Wt %</th>
<th>NH$_4^+$ Wt %</th>
<th>Al$_2$O$_3$/Na$_2$O</th>
<th>Al$_2$O$_3$/CaO</th>
<th>Al$_2$O$_3$/NH$_4$</th>
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</thead>
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<td>NaY</td>
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<td>-</td>
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<tr>
<td>CaY</td>
<td>65.0</td>
<td>22.1</td>
<td>3.5</td>
<td>9.4</td>
<td>-</td>
<td>6.3</td>
<td>2.35</td>
<td>-</td>
</tr>
<tr>
<td>NH$_4$Y</td>
<td>64.8</td>
<td>22.2</td>
<td>3.8</td>
<td>-</td>
<td>9.2</td>
<td>5.84</td>
<td>-</td>
<td>2.41</td>
</tr>
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</table>
Table 3: The chemical analysis of zeolite CaY and NH₄Y by Eisenbach and Gallei (11): (correct version)

<table>
<thead>
<tr>
<th>Sample</th>
<th>SiO₂</th>
<th>Al₂O₃</th>
<th>Na₂O</th>
<th>CaO</th>
<th>NH₄⁺</th>
<th>Al₂O₃/Na₂O</th>
<th>Al₂O₃/CaO</th>
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</tr>
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<td>CaY</td>
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<td>3.6</td>
<td>6.8</td>
<td>-</td>
<td>6.3</td>
<td>3.33</td>
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<tr>
<td>NH₄Y</td>
<td>68.2</td>
<td>23.2</td>
<td>4.0</td>
<td>-</td>
<td>4.6</td>
<td>5.8</td>
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<td>3.04</td>
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Table 4: X-ray data (d, Å') for X and Y Zeolite

<table>
<thead>
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<th>NaY (Å')</th>
<th>NaX (Å')</th>
<th>NaY (Å')</th>
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<td>-</td>
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</table>
There are several possible reasons for the difference between our results and those of Eisenbach and Gallei.

1. The ion exchanges were not carried out in exactly the same way. Possible differences include length of time in solution and temperature, neither of which were specified by Eisenbach and Gallei.

2. Errors in analysis.

3. Zeolites from different sources.

4. Eisenbach and Gallei may have used repeated exchange.

Of these, the most likely are 1 and 4. Despite the difference in degree of ion exchange we felt it worthwhile repeating the experiments carried out by Eisenbach and Gallei in order to determine the influence of the degree of ion exchange on the results. For our experiments we have chosen to use the sample which had the highest degree of exchange (sample 4 of Table 1).

b. The Adsorption:

1. The hydroxyl groups

Fig.1 shows our spectra of CaY zeolite obtained at ambient temperature, 493, 623 and 303 K. A band which occurs at 3690 cm\(^{-1}\) at room temperature (RT) apparently changes its position on heating the sample (493K) to 3640 cm\(^{-1}\), and a broad band at 3560-3100 cm\(^{-1}\) (RT) which gives on heating to 493K a band at 3560-3440 cm\(^{-1}\). The broadness of the band at RT (Fig.1a) probably arises because it consists of two components, a band due to a hydroxyl group (ca.3540 cm\(^{-1}\)) and one due to H\(_2\)O (ca.3400 cm\(^{-1}\)). The latter is mostly removed on heating the sample so that in the higher temperature spectrum (Fig.1b) the band due to the hydroxyl group dominates.
Fig. 1 CaY zeolite at 

a. ambient temperature  
b. 493K  
c. 623K  
d. 303K
Also there is a strong band (Fig.1a) in the deformation region (1650 cm\(^{-1}\)) due to water molecules, which disappears on heating (Fig.1b) and reappears on cooling. After cooling to RT we observe a band at 1695 cm\(^{-1}\) which we assign to the \(v_4\) normal mode of a hydronium ion (Fig.1d). A detailed discussion of the vibrations of the hydronium ion will be given in chapter 5. Eisenbach et al did not report the change in position of the band around 3600 cm\(^{-1}\) on heating. Also, on comparing the intensities (i.e., comparing the heights of the bands not their integrated intensities) of the two corresponding bands in Eisenbach's paper (Fig.2) and in ours (Fig.1), we can see that the band at 3640 cm\(^{-1}\) is more intense than that at 3540 cm\(^{-1}\) (RT) in Eisenbach et al's paper, and becomes stronger on heating to 575 K. In our spectra, however, the bands at 3560-3440 cm\(^{-1}\) are more intense than that at 3640 cm\(^{-1}\) at room temperature (Fig.1b), but they all disappear on heating to 623 K (Fig.1c). In our sample, therefore, all the bands due to the hydroxyl groups disappear on heating to 623 K. This did not happen in the work of Eisenbach and Gallei. The difference in behaviour must stem from the different degree of ion exchange.

The band at 3640 cm\(^{-1}\) (493 K, Fig.1b) has previously been assigned to the hydroxyl stretching vibration in the supercages (17,18) and the band at 3540 cm\(^{-1}\) to the hydroxyl groups in the cubo-octahedra (15-19).

Eisenbach and Gallei found two other bands at, 3750 and 3585 cm\(^{-1}\), as shown in Fig.2a and 2b, which they assign to hydroxyl groups located on the external zeolite crystal surface, and the hydroxyl-stretching vibration of Ca(OH)\(^+\)
Fig. 2a

OH stretching bands of CaY

a. before adsorption of hexene-1
b. after adsorption of hexene-1 in the presence of the gas phase
c. after pumping off the gas phase
d. after pumping off and heating to 570K (ref.1)

Fig. 2b

OH stretching bands of Pt/CaY

a. before adsorption
b. after adsorption of n-hexane and 10 min. of reaction time
c. 22 hr. of reaction time
d. after 22 hr of reaction time followed by evacuation (ref.1)
groups, which are formed by a reaction between adsorbed water molecules and Ca$^{2+}$ ions, respectively. Table 5 summarises our data together with that of Eisenbach et al. Also shown in Table 5, are the assignments of the observed bands (15,16).

**Table 5: Infrared data (cm$^{-1}$) of the OH groups observed by Eisenbach et al, and by us**

<table>
<thead>
<tr>
<th>Eisenbach et al (1)</th>
<th>This work$^a$</th>
<th>assignment (15,16)</th>
</tr>
</thead>
<tbody>
<tr>
<td>RT 493K</td>
<td></td>
<td></td>
</tr>
<tr>
<td>3740</td>
<td>-</td>
<td>external</td>
</tr>
<tr>
<td>3640</td>
<td>3640</td>
<td>supercages</td>
</tr>
<tr>
<td>3585</td>
<td>-</td>
<td>Ca(OH)$^+$</td>
</tr>
<tr>
<td>3540</td>
<td>3560-3100</td>
<td>cubo-octahedra</td>
</tr>
<tr>
<td></td>
<td>3540</td>
<td>H$_2$O str.</td>
</tr>
</tbody>
</table>

$^a$. at 623K all of the bands disappear

Several papers (15-19) have reported these four hydroxyl groups on elevated temperature for CaY zeolite and after dehydration at 743K the bands at 3738, 3640, 3585 and 3540cm$^{-1}$ are still there. Some studies have also reported more than four bands due to OH groups. For NaY zeolite, three bands at 3735, 3635 and 3544cm$^{-1}$ are observed (17). We have been unable to observe the bands assigned by Eisenbach and Gallei to OH groups on the surface of the zeolite and to Ca(OH)$^+$. This could be due to the concentration of these species being very low in our sample.
Eisenbach et al described the importance of each OH group in forming coke. The influence of the external hydroxyl groups (3740 cm\(^{-1}\)) on coke formation is very low, below 773K and the Ca(OH)\(^+\) (3585 cm\(^{-1}\)) groups do not participate at all in the coking reaction. They observed that the highest activity for coking is associated with the hydroxyl groups in the supercages (3640 cm\(^{-1}\)). This band decreases in intensity with increasing intensity of the band at 1585 cm\(^{-1}\) which is due to coke. On the other hand, the band at 3540 cm\(^{-1}\) (cubooctahedra OH) is much less changed in intensity.

In our experiment, it was not possible to see the hydroxyl groups at 3740 and 3585 cm\(^{-1}\) which Eisenbach et al saw at room temperature and on heating the sample, so it is impossible to predict the relation between the hydrocarbon molecules and these hydroxyl groups. These observations may be due to the difference in the degree of ion exchange.

2. Adsorption of the hydrocarbons on CaY zeolite at room temperature

On adsorbing hexene-1 and n-hexane separately on to different discs of CaY zeolite and evacuating for 10 minutes at room temperature several new bands are observed (Figs. 3 and 4). Tables 6 and 7 summarise the frequencies of the bands due to each adsorbed species and the frequencies of the bands in the spectra of the pure zeolite. Three of the bands due to adsorbed n-hexane occur at 2960, 2925 and 2865 cm\(^{-1}\) and hexene-1 at 2956, 2925 and 2850 cm\(^{-1}\) (Tables 6 and 7) which may be assigned to antisymmetric CH\(_3\)-, CH\(_2\)- stretching vibration and symmetric C-H vibrations respectively (19). The bands at 1465 and 1380 cm\(^{-1}\) for n-hexane and
Fig-3

a. hexene-1 adsorbed on CaY zeolite at 293 K

b. hexene-1 liquid
Fig. 4

a. n-hexane adsorbed on CaY zeolite at 293 K
b. n-hexane liquid
Table 6: Infrared data for dehydrated CaY zeolite, CaY - n-hexane and n-hexane liquid (cm$^{-1}$) in the range 4000-1000 cm$^{-1}$ (this work).

<table>
<thead>
<tr>
<th>CaY After baking at 573 K</th>
<th>CaY n-hexane 10 mins evacuation</th>
<th>n-hexane</th>
<th>assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td>$\nu_{\text{as}}$(CH$_3$) (20)</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>$\nu_{\text{as}}$(CH$_2$) (20)</td>
</tr>
<tr>
<td>2910 s</td>
<td>2960</td>
<td>2955 s</td>
<td>$\nu_{\text{as}}$(CH$_3$) (20)</td>
</tr>
<tr>
<td></td>
<td>2925 s</td>
<td>2925 s</td>
<td>$\nu_{\text{as}}$(CH$_2$) (20)</td>
</tr>
<tr>
<td></td>
<td>-</td>
<td>-</td>
<td>instrument background</td>
</tr>
<tr>
<td>2825 m</td>
<td>2865 s</td>
<td>2870, 2855 s</td>
<td>$\nu_{\text{v}}$(CH) (20)</td>
</tr>
<tr>
<td></td>
<td>-</td>
<td>-</td>
<td>instrument background</td>
</tr>
<tr>
<td>1685 m</td>
<td>1710, 1695 s</td>
<td>-</td>
<td>$\nu_4$ (H$_2$O)$^+$</td>
</tr>
<tr>
<td>1650 w</td>
<td>-</td>
<td>-</td>
<td>$\nu_2$ (H$_2$O)</td>
</tr>
<tr>
<td>1470-1380 b</td>
<td>1470 Sh</td>
<td>-</td>
<td>zeolite structure</td>
</tr>
<tr>
<td></td>
<td>1430 w</td>
<td>1465 m</td>
<td>$\delta$ (CH$_3$) (20)</td>
</tr>
<tr>
<td></td>
<td>1380 m</td>
<td>1380 m</td>
<td>$\delta$ (CH$_2$) (20)</td>
</tr>
</tbody>
</table>
Table 7: Infrared data for dehydrated CaY zeolite, CaY-hexene-1, and hexene-1 Liquid (cm$^{-1}$) in the range 4000-1000 cm$^{-1}$ (this work).

<table>
<thead>
<tr>
<th>CaY after baking at 573K</th>
<th>CaY-hexene-1 10 mins evacuation</th>
<th>hexene-1</th>
<th>assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>-</td>
<td>-</td>
<td>3080 s</td>
<td>$\nu$(H-C=C) (20)</td>
</tr>
<tr>
<td>-</td>
<td>2956 w</td>
<td>2960 s</td>
<td>$\nu_{as}$(CH$_3$) (20)</td>
</tr>
<tr>
<td>-</td>
<td>2925 s</td>
<td>2925 s</td>
<td>$\nu_{as}$(CH$_2$) (20)</td>
</tr>
<tr>
<td>2910 s</td>
<td>-</td>
<td>-</td>
<td>instrument background</td>
</tr>
<tr>
<td>-</td>
<td>2850 m</td>
<td>2880,2860 s</td>
<td>$\nu_{s}$(CH) (20)</td>
</tr>
<tr>
<td>2825 m</td>
<td>-</td>
<td>-</td>
<td>instrument background</td>
</tr>
<tr>
<td>1685 m</td>
<td>1710 s</td>
<td>-</td>
<td>$\nu_4$(H$_3$O$^+$)</td>
</tr>
<tr>
<td>1650 w</td>
<td>-</td>
<td>-</td>
<td>$\nu_2$(H$_2$O)</td>
</tr>
<tr>
<td>-</td>
<td>-</td>
<td>1642 s</td>
<td>$\nu$(C=C) (20)</td>
</tr>
<tr>
<td>1470-1380 b</td>
<td>1470 sh</td>
<td></td>
<td>zeolite structure</td>
</tr>
<tr>
<td>-</td>
<td>1450 w</td>
<td>1467,1458 s</td>
<td>$\delta$(CH$_3$) (20)</td>
</tr>
<tr>
<td>-</td>
<td>-</td>
<td>1440,1415 m</td>
<td>$\delta$(CH$_2$) (20)</td>
</tr>
<tr>
<td>-</td>
<td>1370 m</td>
<td>1380 m</td>
<td>$\delta$(CH$_2$) (20)</td>
</tr>
<tr>
<td>-</td>
<td>1235 w</td>
<td>-</td>
<td>$\nu_2$(H$_3$O$^+$)</td>
</tr>
</tbody>
</table>
1450 and 1370 cm$^{-1}$ for hexene-1 are due to the CH$_3$ and CH$_2$ deformation modes (20). No absorption bands from olefinic species at 3080 and 1642 cm$^{-1}$ of hexene-1 adsorbed on to CaY were detected after 10 minutes evacuation (Table 7). This is in agreement with the data of Eisenbach and Gallei, who explained this by proposing a linear adsorbed hydrocarbon species formed by interaction of the strong acidic hydroxyl groups with the π electrons of hexene-1 according to the equation:

\[
\text{CH}_2 = \text{CH} = \text{CH}_2 + \text{OH} \rightarrow \text{CH}_2 = \text{CH} = \text{CH}_2 + \text{OH} < \text{CH}_2 = \text{CH}_2 <\text{CH}_2 = \text{CH}_2 (1)
\]

In addition to the above results, we can see from Tables 6 and 7 that there are two bands at 1710 and 1240 cm$^{-1}$ which were not found in the work of Eisenbach and Gallei. A band at 1710 cm$^{-1}$ was observed with both hexene-1 and n-hexane but the band at 1240 cm$^{-1}$ was observed only for hexene-1 (Fig.3). The 1240 cm$^{-1}$ band could not be seen with n-hexane and this may be because:

1. There is a very intense absorption due to the framework in the region 1250-900 cm$^{-1}$ which may conceal the band expected at ca 1240 cm$^{-1}$ for n-hexane, or

2. If the bands at ca 1710 and 1240 cm$^{-1}$ are due to the same species then since the 1240 cm$^{-1}$ band is very much less intense than that at 1710, it will be more difficult to observe in the spectrum of adsorbed n-hexane than hexene-1, since the intensity of the 1710 cm$^{-1}$ band is very much less in the former (Fig.4) than in the latter case (Fig.3).
These two bands can be assigned to the hydronium ion which is present because there are some water molecules in the zeolite as was seen on cooling the sample (Fig.1d). In the case of hexene-1 these water molecules could interact with the proton released in the formation of the adsorbed species (equation 1), to form the hydronium ion. The band in the region of 3100-2800cm$^{-1}$, which is due to the hydronium ion, could not be seen because of the presence of three bands due to the instrument background in the same region. The presence of the hydronium ion in zeolites will be discussed in more detail later (chapter 5).

3. The effect of temperature on adsorbed hydrocarbons

The adsorption of the hydrocarbons on CaY zeolite at room temperature followed by heating to higher temperature in the presence of the gas phase causes dramatic changes to be observed in the infrared spectra as a function of temperature. Since excess hexene-1 is present, the spectra shown in Fig.5 are a mixture of gas phase and adsorbed species. Fig.5 shows that as the temperature is increased hexene-1 loses its double bond character as evidenced by the absence of an absorption band near 3080cm$^{-1}$ due to $\nu (H-C = C)$. The spectra show that the gas phase hexene is reacting with the zeolite and being consumed. New bands at 1580 and 1340cm$^{-1}$ appear at high temperature. These two bands become more intense as the band due to $\nu (H-C = C)$ decreases in intensity and as the temperature is increased (Fig.5). The band at 1340cm$^{-1}$ is assigned to a vibration of tertiary C-H groups (20).
4. Coke formation

When either hexene-1 or n-hexane is added to CaY zeolite at room temperature and the sample heated, a new band appears at $1580\text{cm}^{-1}$. From Figs. 5 and 6 we can see that the intensity of this band increases with increasing temperature.

Infrared studies by Unger and Gallei (21) of activated carbon and soot samples show a band at $1585\text{cm}^{-1}$ which was assigned to the $\text{C=C}$ stretching vibration of microcrystalline graphitic carbon structures, which are present in polycyclic aromatic compounds.

The coke formation at higher temperatures can be seen only if the gas phase is present over the zeolite surface. Also coke formation will be effected by the type of hydrocarbon which is in contact with the zeolite surface. When we compare the two hydrocarbons, hexene-1 and n-hexane it can be seen that hexene-1 can form coke faster than n-hexane (Fig.6), and confirms the observations of Eisenbach and Gallei. The band due to coke is very much stronger in the case of hexene-1 than for n-hexane (Fig.6), indicating that coke is formed much more easily from the unsaturated hydrocarbon. Radioactive tracer studies conducted by Hightower and Emmett (22) have indicated that the olefins have the highest ability to form coke. Also it was found (2) that the saturated hydrocarbons are very much less reactive than their corresponding olefins.

Eisenbach and Gallei have assumed that the hydroxyl groups react with adsorbed hydrocarbon species producing coke and thereby are consumed irreversibly. We can see,
Fig. 6  
a. n-hexane adsorbed on CaY zeolite at 573K  
b. hexene-1 adsorbed on CaY zeolite at 573K
(Fig. 5) that we have observed two new bands at 3740 and 3600cm$^{-1}$ which appeared when CaY zeolite was heated with hexene-1 at 550K. Since these hydroxyl groups appear to be formed only on the breakdown of the hexene-1 the group giving a band at 3740cm$^{-1}$ is unlikely to be the same as that observed by Eisenbach and Gallei at room temperature at the same frequency.

c. Conclusion

Several of our observations are in agreement with those of Eisenbach and Gallei:

1. The appearance of a "coke-band" for both hexene-1 and n-hexane adsorbed on CaY zeolite.

2. The unsaturated hydrocarbon forms coke more readily than the saturated one.

3. There are several bands due to the species adsorbed at room temperature.

4. In the case of hexene-1 the adsorbed species does not have any olefinic character.

5. We observe bands due to O-H groups at 3640 and 3540cm$^{-1}$.

In some ways, however, our results differ from theirs:

1. We observe bands due only to two types of hydroxyl groups while Eisenbach and Gallei observe four at RT.

2. At high temperature we could not observe any bands due to hydroxyl groups and so could not correlate coking activity with a particular hydroxyl site as Eisenbach and Gallei were able to do.

3. Two new hydroxyl bands appear at high temperature in the case of hexene-1.
References


11. Gallei, E., Privet communication


CHAPTER 5

INFRARED STUDIES OF WATER ADSORBED ON
PARTIALLY MANGANESE OR COBALT EXCHANGED
ZEOLITES

1. Introduction

Most spectroscopic studies of adsorption on zeolites so far reported discuss the spectra of water molecules and hydroxyl groups (1-22). The infrared spectra of water adsorbed on several synthetic zeolites has been studied by a number of experimentalists (19-23). Of particular relevance to this work the influence of dehydration and adsorption of water on

1. the coordination of the cations,
2. the state of the adsorbed water,
3. the stretching vibrations of the water molecule,

in CoNaA and NaA zeolite has been investigated by means of electronic and infrared spectroscopy (16).

The presence of the hydronium ions which compensate for cation deficiencies in zeolites has been discussed (24) and the bands corresponding to the normal modes of the hydronium ion reported (25). The existence of the hydronium ion as a discrete chemical entity has been known since the early 1920s. From nuclear magnetic resonance studies Richards and Smith (25) were able to measure H-H separations (~1.72Å) in the hydronium ion in $\text{H}_3\text{O}^+\text{ClO}_4^-$. They concluded that the hydronium ion was a nearly flat pyramidal species with O-H bond lengths of ~1.02Å and H-O-H angles of 115°.
Neutron diffraction studies indicated a pyramidal $C_{3v}$ symmetry for the hydronium ions with O-H bond length of $\sim 0.98\text{Å}$ and H-O-H angle of $111^\circ$. A spectroscopic study and normal coordinate analysis of the $H_3O^+$ and $D_3O^+$ species in solid $H_3O^+\text{CH}_3\text{C}_6\text{H}_4\text{SO}_3^-$ were made (30).

In Chapter 2, a general account of the structure of A type zeolites was given. A more detailed account of the Mn-4A and Co-4A zeolites structures will be given in this chapter.

a. Structure of partially exchanged Mn-4A zeolite.

1. Dehydrated

An X-ray study of dehydrated partially Mn-exchanged A zeolite (31-33) revealed a total of 4.5 Mn(II) and 3 Na$^+$ ions per supercage and these were located near the centres of 6-rings. The Mn(II) ions are recessed 0.108Å into the sodalite unit from the plane of three oxygens O(3) (Fig.1a,b). The Na$^+$ ions are 0.46Å from the corresponding plane but are recessed into the large cavity. Both Mn(II) and Na$^+$ ions are trigonally coordinated to respective sets of three framework O(3)'s at 2.11 and 2.16Å respectively.

2. Hydrated

In hydrated MnA, each of the 4.5 Mn(II) ions lies in a 6-ring where the two Mn(II)-OH$_2^-$ distance are 2.03 and 2.06Å for H$_2$O(1) and H$_2$O(2) (Fig.2a,b). The three symmetry-equivalent Mn(II) to framework O(3) distances are 2.28Å. These complete a slightly distorted bipyramid with Mn displaced 0.2Å into the larger cage from the plane of the O(3) ions.
Fig. 1a

Fig. 2a

Fig. 1b

Fig. 2b
b. Structure of partially exchanged Co-A zeolite

1. Dehydrated

Dehydrated partially Co-exchanged A zeolite (34,35) has a structure similar to the manganese form with the Co(II) ions occupying threefold axis sites near the 6-ring oxygen windows, but recessed by 0.16Å into the sodalite cavity (Fig.3). Each Co(II) is 2.08Å from three equivalent O(3) ions arranged trigonally in the six-oxygen ring, and 4 Na⁺ ions are at 2.12Å from the O(3) ions.

2. Hydrated

Amaro et al (35) have reported data for hydrated Co₄Na₄-A zeolite. One Co(II) ion lies at the centre of the sodalite cage where it is surrounded by 6 water molecules (Co-OH₂=2.11Å) each hydrogen-bonded to two framework oxygens. Three Co(II) ions lie on or near a triad axis displaced 1.6Å into the larger cage from the plane of a 6-ring (Fig.4). The distance to the nearest framework oxygens is 2.7Å. Each Co(II) is tetrahedrally coordinated to one H₂O(2) and three oxygens (36).

2. Experimental

Materials

Partially exchanged Mn-4A and Co-4A zeolites were prepared from NaA zeolite (BDH Chemicals Ltd) by ion exchange with 0.1N MnSO₄ and Co(NO₃)₂ (BDH analar grade) solutions respectively following the method described by Seff (37). Two additional different Mn concentrations were used to obtain different degrees of ion exchange as shown in Table 1.
Thermogravometric analysis of each Mn-4A and Co-4A zeolites were obtained.

Procedure
Self supporting discs were evacuated in the infrared cell, then heated under vacuum to 533K and cooled to the desired temperature. Water was admitted to the cell, the sample cooled to room temperature and evacuated then heated to 533K. Spectra were obtained at all stages.

The exchange of D$_2$O for H$_2$O

Exchange of D$_2$O for H$_2$O in Mn-4A zeolite (which had previously been dried at 323K for one day) was attempted using several different methods.

1. 2g of Mn-4A zeolite was dispersed in 10 mls of D$_2$O and stirred at room temperature for one week. The sample was dried under vacuum. A similar procedure was used involving 30 ml of D$_2$O with 1g of Mn-4A zeolite but in this case the temperature of the sample was increased to 363K in a nitrogen atmosphere for 5 days.

2. In the second method 1g of Mn-4A zeolite was heated under vacuum over a period of 17 hours, to 673K. The sample was then cooled slowly to 573K when D$_2$O was admitted and the sample cooled to room temperature.

3. In the third method a Carius tube was used with 20 ml of D$_2$O and 1g of Mn-4A zeolite. The tube was evacuated and sealed then left in a furnace for two days at 473K, the sample was dried under vacuum.
Table 1: Chemical analysis of the Mn-4A and Co-4A zeolites (% by weight)

<table>
<thead>
<tr>
<th></th>
<th>N</th>
<th>Al</th>
<th>Na</th>
<th>Mn</th>
<th>Co</th>
<th>Mn(II)/Na(I)</th>
<th>Co(II)Na(I)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Mn-4A</td>
<td>0.03</td>
<td>12.6</td>
<td>6.65</td>
<td>2.95</td>
<td>-</td>
<td>1.061</td>
<td>-</td>
</tr>
<tr>
<td>Mn-4A</td>
<td>0.07</td>
<td>14.79</td>
<td>4.37</td>
<td>7.71</td>
<td>-</td>
<td>4.219</td>
<td>-</td>
</tr>
<tr>
<td>Mn-4A</td>
<td>0.1</td>
<td>12.77</td>
<td>3.44</td>
<td>7.41</td>
<td>-</td>
<td>5.151</td>
<td>-</td>
</tr>
<tr>
<td>Co-4A</td>
<td>0.1</td>
<td>12.9</td>
<td>0.75</td>
<td>-</td>
<td>15.55</td>
<td>-</td>
<td>53.185</td>
</tr>
</tbody>
</table>

a. normality
4. A disc of \( \sim 10\text{mg} \) of Mn-4A zeolite was inserted in an infrared cell then heated under vacuum to 553K. The sample was cooled to 473K and D\(_2\)O inserted, then it was cooled to room temperature. Spectra were obtained at all stages.

3. Results and discussion:

The spectrum of Mn-4A zeolite (Mn(II)/Na\(^+\) ratio = 5.15) before heating (Fig.5) shows a band at 1355cm\(^{-1}\). This band disappears on heating (Fig.5c). To determine whether or not this band was associated with the water molecules, it was decided to exchange D\(_2\)O for H\(_2\)O since if the band was due to water it should shift to lower frequency on adsorption of D\(_2\)O.

a. Ion and molecular exchange

(1) D\(_2\)O exchange

From Fig.6 we can see that it is difficult to exchange the H\(_2\)O molecules inside the zeolite. None of the first three methods gave any sign of the presence of D\(_2\)O exchange except for a small band which occurred after using the Carius tube method (Fig.6b). The spectra do not show any sign of the bands (2600 and 1210cm\(^{-1}\)) which are characteristic of D\(_2\)O. In the spectrum of Mn-4A-H\(_2\)O we can see that there are two bands which can be easily assigned to vibrations of the H\(_2\)O molecules. The first at 3500cm\(^{-1}\) which is very broad, and the second is at 1650cm\(^{-1}\). On exchanging D\(_2\)O for H\(_2\)O it is expected that these two bands will shift to lower frequency and occur at \( \sim 2600 \) and
1210 cm\(^{-1}\) respectively. In Fig.6a, however, we can see only those bands due to H\(_2\)O molecules and there is no sign of the bands one would expect if D\(_2\)O had been exchanged for H\(_2\)O in the cages. In Fig.6b, while the bands due to H\(_2\)O dominate the spectrum there is a weak band at 2600 cm\(^{-1}\) which is due to D\(_2\)O. Therefore even under the stringent conditions of the Carius tube experiment, only a small degree of exchange has occurred. These results are indicative of the difficulties of exchanging H\(_2\)O for D\(_2\)O.

When the experiments were first attempted, it was expected that the replacement of H\(_2\)O by D\(_2\)O molecules would be relatively easy because the water molecules are regarded as being mobile within the framework. Several studies (20,21) on D\(_2\)O exchange in zeolite, by adsorption of D\(_2\)O on a dehydrated zeolite disc which gives bands due to the D\(_2\)O molecules have been reported. Our experiments using this technique were also successful (see later discussion).

This difficulty of exchanging the H\(_2\)O molecules inside the zeolite is rather surprising, and it may mean that the D\(_2\)O is not able to enter the same sites as the H\(_2\)O in a hydrated framework. It was found (22) that surface OH groups in Na and NH\(_4\) X zeolite do not exchange rapidly with the physically sorbed deuterium oxide. So Lynch et al (22) saturated the zeolite with the heavy water and both the OH and OD frequencies were observed. No further experimental details were given.

Since type A zeolites have smaller pores than X zeolites, it will be expected that if zeolite X does not exchange rapidly with deuterium oxide (22), we can expect that A zeolite be even more difficult.
(2) **Ion exchange**

We will discuss here the spectral region 1600-1300cm\(^{-1}\). This region contains at low degree of exchange (Mn(II)/Na\(^+\) ratio = 1.06) one broad band at 1440-1350cm\(^{-1}\) (Fig. 7a) at room temperature. This band does not change its position on heating. On comparing this spectrum with that of NaA zeolite, we can see that there is a band at 1425cm\(^{-1}\) (Table 2) in the later case. Table 2 shows also that there is a band at 1355cm\(^{-1}\) which is present in zeolites which have a high degree of Mn exchange (Mn(II)/Na\(^+\) ratio = 5.15) and could be due to the presence of Mn ions. From these two observations there is no alternative but to assign the broad band at 1440-1350cm\(^{-1}\) observed in the sample with low degree of exchange, to two vibrations associated with the presence of two types of cations (Mn(II) and Na\(^+\)).

**Table 2: Infrared data (cm\(^{-1}\)) of Mn-4A zeolite at different degree of exchange and of NaA zeolite (1600-1300cm\(^{-1}\))**

<table>
<thead>
<tr>
<th>Temperature (K)</th>
<th>Mn-4A zeolite (Mn(II)/Na(^+) ratio = 1.06)</th>
<th>Mn-4A zeolite (Mn(II)/Na(^+) ratio = 5.15)</th>
<th>NaA zeolite</th>
</tr>
</thead>
<tbody>
<tr>
<td>R.T. 1440-1350s-b</td>
<td>1355 s</td>
<td>1395 w-sh</td>
<td>1375 m</td>
</tr>
<tr>
<td>573 1440-1350s-b</td>
<td>1395 s</td>
<td>1540 s</td>
<td>1595-1565 w-b</td>
</tr>
</tbody>
</table>

s=stronge
b=broad
m=medium
w=weak
Fig. 7  \( \text{Mn}_{1.064} \text{A zeolite.} \) a. at 293K
b. at 573K
c. after adding \( \text{H}_2\text{O} \) at 473K
d. after adding \( \text{H}_2\text{O} \) at 573K
Riley and Seff (32,33 and 38) have suggested that dehydrating Mn-4A zeolite (Mn(II)/Na\(^+\) ratio = 1.5) causes the Mn ions to move from a position slightly inside the sodalite unit to one slightly inside the large cage. They therefore found two sites, one site present on the hydrated form and the other on the dehydrated. In our spectra, however, we could not see any change in the position of the band at 1440-1350 cm\(^{-1}\) on heating.

At higher degree of Mn-exchange (Mn(II)/Na\(^+\) ratio = 5.15) it can be seen that a strong band appears at 1355 cm\(^{-1}\) immediately after recording the spectrum (Fig.5a); also a weak band at 1395 cm\(^{-1}\) appears after evacuation for one night. This implies that the Mn ions occupy a site at room temperature which is associated with the presence of the strong band at 1355 cm\(^{-1}\). At the same time if we look to the spectra we can see that with NaA zeolite there is a band at 1425 cm\(^{-1}\) which is due to the presence of Na ions. MnA zeolite with low degree of exchange shows a broad band at 1440-1350 cm\(^{-1}\) which is assigned to two vibrations, each one associated with the presence of the Na\(^+\) ions, and Mn ions. The MnA zeolite spectrum of the high degree of exchange sample has a strong band at 1355 cm\(^{-1}\) assigned to vibration associated with the presence of Mn ions and a weak band at 1395 cm\(^{-1}\) associated with the presence of Na ions. The observed change in relative intensity of the bands associated with the Na\(^+\) and Mn(II) ions is in agreement with the alteration in the degree of exchange. This can be explained by the movement of Mn(II) ions from one site to a different one. The Mn ions, when they move coincidentally, have the same frequency as that associated
with the Na ions (1395cm\(^{-1}\)). This band will therefore coincide with the small band which is due to the presence of Na\(^+\) ions. From these facts one can assign the two bands at 1355 and 1395cm\(^{-1}\) to the Mn ions in different sites which are present at different temperatures.

A band at 1535cm\(^{-1}\) has been observed at room temperature with high Mn exchange and does not appear with low degree of exchange (Fig.5). This band does not change on heating which implies that a totally different site has been occupied with Mn ions when a high degree of Mn has been exchanged.

From the above one can conclude that at high degree of exchange the Mn ions will occupy two sites at room temperature and two sites at high temperature.

b. The hydroxy1 groups

Bands at 3740 and 3640cm\(^{-1}\) were observed in the spectra of Co-4A and Mn-4A(Mn(II)/Na\(^+\) ratio = 5.15) zeolites (Figs. 5 and 8) respectively. The weak band at 3740cm\(^{-1}\) observed at 348K in Co-4A zeolite is still present at 548K while the band observed at 3640cm\(^{-1}\) in Mn-4A zeolite disappears at 423K. These two bands have been observed in several studies (39-41), with a variety of cations in A, X and Y zeolites. The band at 3740cm\(^{-1}\) has been assigned to the Si-OH groups which terminate the external surface and this corresponds with the observations for CaY zeolite (40). The band at 3640cm\(^{-1}\) has been assigned to interaction of the hydroxyl groups with the cations, since this band has been observed to shift with different cations (39).
The broad band at 3700-3300 cm\(^{-1}\) (Fig. 8) can be assigned to water molecules which might form a hydrogen bond with a lattice oxygen and this may be the reason for the non-disappearance of this band even when heated to 558 K. From the TGA (Fig. 9) of Mn-4A zeolite (Mn(II)/Na\(^+\) ratio = 5.15) and Co-4A zeolite, there is a big step at around 413 K which is due to the loss of water molecules. With increasing temperature there is a continuous loss of the water molecules until 673 K. The water being lost is that which gives rise to the band at 3600-3300 cm\(^{-1}\).

Strong absorption bands at 1655 and 1645 cm\(^{-1}\) were observed in the spectra of Mn-4A and Co-4A zeolites respectively (Figs. 5a and 8) and these are completely removed after evacuation of these zeolites at approximately 473 K. These bands are due to deformation vibrations of the water molecules (39, 40); and arise from those molecules which do not form a hydrogen bond and which are therefore probably interacting with the cation via the oxygen atom.

c. The presence of the hydronium ion

On adding water to the Mn-4A zeolite (Mn(II)/Na\(^+\) ratio = 5.15) at approximately 403 K and then cooling, a broad band at 3700-3000 cm\(^{-1}\) was observed (Fig. 10). This is due to the presence of excess water molecules. As expected from this, the band at 1655 cm\(^{-1}\) can also be observed. These observations correspond with those already reported for adsorbed water (40). There is however a new band at 1685 cm\(^{-1}\) (Fig. 10) which only appears when water vapour is admitted to a hot zeolite. Table 3 summarises our results for Mn-4A zeolite.
Fig. 9  thermogravometric analysis of Co-4A and Mn-4A zeolite
before and after adding water vapour.

Table 3: Infrared data (cm\(^{-1}\)) for Mn-4A

\[(\text{Mn}(\text{II})/\text{Na}^+ \text{ ratio} = 5.15) \text{ zeolite} \]

and of adsorbed water in the range 4000-1000cm\(^{-1}\).

<table>
<thead>
<tr>
<th></th>
<th>before adding water</th>
<th>after adding water at 403K</th>
<th>after adding water at 293K</th>
</tr>
</thead>
<tbody>
<tr>
<td>293K</td>
<td>533K</td>
<td>403K</td>
<td>293K</td>
</tr>
<tr>
<td>3700-3000cm(^{-1})</td>
<td>3600-3400cm(^{-1})</td>
<td>3600-3300cm(^{-1})</td>
<td>3600-2900cm(^{-1})</td>
</tr>
<tr>
<td>1655 s</td>
<td>-</td>
<td>1685 s</td>
<td>1695 s</td>
</tr>
<tr>
<td>1535 s</td>
<td>-</td>
<td>1555 m</td>
<td>1520 sh</td>
</tr>
<tr>
<td>1355 s</td>
<td>-</td>
<td>1365 s</td>
<td>1370 s</td>
</tr>
</tbody>
</table>

Assignment:
- \(v_1, v_3\) (OH)
- \(v_4\) (\(\text{H}_3\text{O}^+\))
- \(v_2\) (\(\text{H}_2\text{O}\))
- Zeolite structure

As the zeolite was heated, the water band at 1655cm\(^{-1}\)
disappeared while the band at 1685cm\(^{-1}\) remained (Fig.11).
A band at 1230cm\(^{-1}\) also appeared upon adding water vapour
to the hot zeolite and became more intense at room temperature
(Fig.10). On heating the sample to 573K the band at 1230cm\(^{-1}\)
slowly disappeared also (Fig.11). At the same time the sharp
band at 1685cm\(^{-1}\) decreases slightly in its intensity. When
we compare these spectra (Figs.11a,b) with the spectrum of
liquid water (Fig.11c) [see also ref.(26,42-45)], we can
deduce that the bands in the region 1300-1200cm\(^{-1}\) cannot
be due to the water molecules. The bands found in the
Fig. 10

Mn-4A zeolite a. at 573K before adsorbing H₂O
b. at 405K after adding H₂O
c. at 293K
spectrum of the zeolite at 1685 and 1230 cm\(^{-1}\) must therefore be due to other species. The species most likely to be present is the hydronium ion. From the infrared spectrum of the hydronium ion in crystals (26, 46) we find that the vibrational frequencies of the \(\text{H}_3\text{O}^+\) and \(\text{D}_3\text{O}^+\) ions are as shown in Table 4.

**Table 4: The infrared data (cm\(^{-1}\)) for \(\text{H}_3\text{O}^+\) and \(\text{D}_3\text{O}^+\) ions in crystals (26, 46)**

<table>
<thead>
<tr>
<th>(\text{H}_3\text{O}^+ \text{Cl}^-)</th>
<th>(\text{D}_3\text{O}^+ \text{Cl}^-)</th>
<th>assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>1150</td>
<td>785</td>
<td>(\nu_2)</td>
</tr>
<tr>
<td>1700</td>
<td>1255</td>
<td>(\nu_4)</td>
</tr>
<tr>
<td>2540-3235</td>
<td>2000-2445</td>
<td>(\nu_3, \nu_1)</td>
</tr>
</tbody>
</table>

The normal modes of vibrations for \(\text{H}_3\text{O}^+\) are shown in Fig.12.

The hydronium ion in NaX and NH\(_4\)X zeolites have been shown to be unstable and to decomposes when the zeolites are heated in vacuo (47). Szymanski et al (21) have also reported bands due to hydronium ion on NaX zeolite at 1700-1750 cm\(^{-1}\), at high water concentration.

From the above we can assign the bands at 1685 and 1230 cm\(^{-1}\) to the hydronium ion. We must also consider the fact that while the band at 1230 cm\(^{-1}\) disappears at 573K the band at 1685 cm\(^{-1}\) only decreases slightly in intensity (Fig.11).
One possible explanation for this is that when the hydronium ion decomposes, the water molecules which are formed, form strong hydrogen bonds with the zeolite framework via the interaction of the framework oxygen atoms with the hydrogen atom of the water molecules (39, 48, 49). This is indicated by the fact that the band at 1685\text{cm}^{-1} is not removed even after evacuation at 573K. This band we assign to the HOH bend of the hydrogen bonded water molecules. If this assignment is correct then as we observe the band at 1685\text{cm}^{-1} we also expect to observe a band in the O-H stretching region due to the presence of the hydrogen bond. Because the H$_2$O is strongly hydrogen bonded the O-H stretch should be reduced considerably in frequency relative to the value in liquid water. There is no suitable band in the spectrum and it is possible that it is too weak to observe.

It should be noted that the stretching mode ($\nu_1, \nu_3$) of the hydronium ion (observed in H$_3$O$^+$Cl$^-$ crystal at 2540-3235 cm$^{-1}$) was also not observed in our spectrum (Fig.10). The stretching mode of the hydronium ion known to be weak and broad relative to $\nu_2, \nu_4$ and may explain why it was not observed in our spectra since $\nu_2$ was not itself very intense. It is also possible that the band at
Fig. 12 Normal modes of vibration for $\text{H}_2^+$ (30)
$1230\text{cm}^{-1}$ does not disappear on heating but that the broad zeolite absorption ($1250-900\text{cm}^{-1}$) increases in width and the two are not resolved.

When the same procedure was carried out with samples of different degrees of exchange of Mn(II) for Na$^+$ (Table 1), the intensity of the band at $1685\text{cm}^{-1}$, which we have assigned to the hydronium ion, varies with the degree of ion exchange. From Figs. 7, 10 and 11, it can be seen that the intensity of this band is strong with the zeolite of high Mn content ($\text{Mn(II)}/\text{Na}^+$ ratio = 5.15) and is less intense with the Mn(II)/Na$^+$ ratio = 1.06 (Fig. 7). This suggests that the hydronium ions are bonded to the Mn ions rather than Na$^+$ ions.

In the case of the Co-4A zeolite, when water vapour is added a broad band occurs at $3700-3000\text{cm}^{-1}$ (Fig.13), which is due to the presence of water, and a band at $1645\text{cm}^{-1}$ due to $v_2(\text{H}_2\text{O})$. There are no bands at $1685$ or $1230\text{cm}^{-1}$, however, which can be assigned to the hydronium ion. The hydronium ion is therefore not formed in this case. A study (16) on the interaction of CoNaA zeolite with water has been reported in which it was noticed that on hydration and dehydration there was no water dissociation. From their spectrum we can see that hydronium ions were not formed in that case either. From this we can conclude
that the hydronium ion is formed in the Mn-4A zeolite but not in Co-4A zeolite.

When doing the D$_2$O exchange experiments only the experiment of admitting D$_2$O vapour to Mn-4A zeolite (Mn(II)/Na$^+$ ratio = 5.15) was successful. The spectrum (Fig.14) shows a broad band at 2700-2100cm$^{-1}$ and this band is due to $\nu_1$ and $\nu_3$ of D$_2$O molecules. In this case it is difficult to see the bend of the D$_2$O molecules since it occurs at $\sim$1178cm$^{-1}$ in liquid D$_2$O and hence is amongst the strong framework absorption below 1250cm$^{-1}$. For the same reasons, we cannot tell whether D$_3$O$^+$ is present since we can see from Table 4 that the two bands corresponding to those we observed for H$_3$O$^+$ would occur beneath the strong framework bands (Fig.14). The infrared data for Mn-4A zeolite and D$_2$O adsorbed at 527K is shown in Table 5.

Table 5: Infrared data (cm$^{-1}$) of Mn-4A with water adsorbed at 527K (3000-1000-cm$^{-1}$) the spectra recorded at different temperatures.

<table>
<thead>
<tr>
<th></th>
<th>527K</th>
<th>289K</th>
<th>513K</th>
</tr>
</thead>
<tbody>
<tr>
<td>2700-2500 b-w</td>
<td>2700-2300 b-s</td>
<td>2600-2400 b-m</td>
<td>assignment</td>
</tr>
<tr>
<td>1655 w</td>
<td>1645 m</td>
<td>-</td>
<td>$\nu_1$ $\nu_3$ (D$_2$O)</td>
</tr>
<tr>
<td>1540 s</td>
<td>1540 m</td>
<td>1540 s</td>
<td>$\nu_2$ (H$_2$O)</td>
</tr>
<tr>
<td>-</td>
<td>1440 m</td>
<td>-</td>
<td>$\nu_2$ (HDO)</td>
</tr>
<tr>
<td>1395 s</td>
<td>1380 m</td>
<td>1395 s</td>
<td>zeolite structure</td>
</tr>
</tbody>
</table>

zeolite structure
Fig. 13  

a. Zeolite after adsorbing $\text{H}_2\text{O}$ at 463K 

b. Zeolite after adsorbing $\text{H}_2\text{O}$ at 463K then cooling to 293K
Fig. 14  

a. Mn-4A zeolite-D$_2$O at 473 K  
b. Mn-4A zeolite-D$_2$O at 320 K  
c. D$_2$O
From Table 5 it can be seen that a band at 1645 cm\(^{-1}\) appears when \(D_2O\) is added which is due to the presence of residual \(H_2O\). Also in the presence of residual \(H_2O\) a new band occurs when adding \(D_2O\) vapour to a hot zeolite, at 1440 cm\(^{-1}\) (Fig. 14), which can be assigned to HDO (49) as shown in Table 6.

Table 6: Infrared data (cm\(^{-1}\)) of the \(D_2O\) and HDO vapour and their assignment (49)

<table>
<thead>
<tr>
<th></th>
<th>(D_2O)</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>HDO</td>
<td>1402</td>
<td>(\nu_2)</td>
</tr>
<tr>
<td></td>
<td>2719</td>
<td>(\nu_1)</td>
</tr>
<tr>
<td></td>
<td>2809</td>
<td>(\nu_3)</td>
</tr>
<tr>
<td></td>
<td>1178</td>
<td>(\nu_2)</td>
</tr>
<tr>
<td></td>
<td>2789</td>
<td>(\nu_1)</td>
</tr>
</tbody>
</table>

4. Conclusion:

A. The Mn ions occupy two different sites in the presence of high degree of Mn ion exchange. On heating the Mn ions will leave one of these sites and transfer to a different one, which will give a total of three sites which might be occupied by the Mn at different temperatures.

B. Hydronium ions are only formed on adding water to a hot zeolite.

C. The presence of the hydronium ion depends on

(1) The degree of ion exchange, since with high degree of Mn ions a strong band will occur due to the
hydronium ion and at low degree of Mn ions give weak band due to $H_3O^+$

(2) The cations present, since Co(II) does not give any bands due to the hydronium ions.

D. It has not proved possible to exchange $D_2O$ for $H_2O$ in A zeolite at room temperature.
References


Chapter 6

An infrared study of cyclopropane adsorbed on Co(II) and Mn(II) partially exchanged A zeolites

1. Introduction

The isomerization of cyclopropane to propylene is a widely used test reaction for the investigation of the performance of catalytic reactors (1) and the activity of acid-type solid catalysts (2). Bassett and Habgood (1) and Habgood and George (3) have studied the isomerization on various cation exchanged forms of zeolite Y at a temperature of 200°C. Bartley et al (4) and George and Habgood (5) in their studies using deuterated zeolite Y have concluded that both exchange and isomerization of cyclopropane proceed via a non-classical protonated cyclic carbonium ion intermediate (6-10).

Infrared spectroscopy was used to study the chemisorption and reactions of cyclopropane over zeolite HY at room temperature by Tam et al (11). Isobutane was the major product. A mechanism for this transformation involving the formation of a non-classical protonated cyclopropane ion intermediate was proposed. Other studies have been made to investigate the skeletal isomerization of cyclopropane over NaY, NaCaY and NaHY zeolites (11-14).

Structure of cyclopropane adsorbed on Mn(II) and Co(II) Exchanged A zeolite

Seff et al (15) determined the crystal structures of the cyclopropane sorption complexes of partially Co(II)-exchanged and partially Mn(II)-exchanged zeolite A (Figs.
1 and 2), the unit cell constants are
12.147 Å for \( \text{Co}_4\text{Na}_4\text{Si}_{12}\text{Al}_{12}\text{O}_{48}.4\text{C}_3\text{H}_6 \) and 12.1466 Å for
\( \text{Mn}_4\text{Na}_4\text{Si}_{12}\text{Al}_{12}\text{O}_{48}.4\text{C}_3\text{H}_6 \).

In each structure transition metal cations are located on
three-fold axes inside the large cavity, close to three
trigonally arranged framework oxide ions on alternate
6-oxygen rings. The metal to oxygen distances are
\( \text{Co}(\text{II})-\text{O} = 2.1748 \) and \( \text{Mn}(\text{II})-\text{O} = 2.1228 \). There are four
cyclopropane molecules per unit cell, each of them is found
to complex to a transition metal ion with \( \text{Co}(\text{II})-\text{C}=2.818 \)
and \( \text{Mn}(\text{II})-\text{C}=3.098 \). No interaction between \( \text{Na}^+ \) ions and
cyclopropane molecules was observed. The adsorption complexes
represent novel chemical species since stable organometallic
complexes of \( \text{C}_3\text{H}_6 \) have not been reported in the scientific
literature.

2. **Experimental**

**Materials**

\( \text{Mn-Na-4A} \) and \( \text{Co-Na-4A} \) zeolites were prepared from
\( \text{NaA} \) zeolite (BDH Chemicals Ltd) by ion exchange with
0.1N \( \text{MnSO}_4 \cdot 4\text{H}_2\text{O} \) and \( \text{Co(NO}_3\text{)}_2 \) (BDH analar grade) solutions
respectively (\( \text{Mn}(\text{II})/\text{Na}^+ \) ratio = 5.15 and \( \text{Co}(\text{II})/\text{Na}^+ \) ratio = 53.18 )
following the method described by Seff (15). These samples
were also used for part of work in the previous chapter
(chapter 5). Cyclopropane gas of 99.8% was used without
further purification.

**Procedure**

The self-supporting disc was evacuated in the infrared
cell, then heated under vacuum and cooled to room temperature.
Fig. 1a A stereoview of the Co$_4$A$_4$-4C$_3$H$_6$ unit cell

Fig. 1b

The coordination environment of the Co(II) ion.
Fig. 2a A stereoview of the Mn₄A₄C₃H₆ unit cell

Fig. 2b

The coordination environment of the Mn(II) ion.
Cyclopropane gas was admitted to the cell which then was evacuated for 10 minutes before obtaining the spectra. A spectrum of cyclopropane gas was also obtained.

3. Results and discussion
   a. The hydroxyl groups
      These were discussed in chapter 5.
   b. The adsorption of cyclopropane
      Figs. 3 and 4 show the spectra of Mn.4A and Co-4A zeolite (after baking and cooling to room temperature) before and after adsorbing cyclopropane. An isolated cyclopropane molecule has $D_{3h}$ symmetry. The activities of the normal modes are summarized in Table 1 together with their symmetry classes (16,17). The spectra of cyclopropane gas is shown in Figs. 3c and 4c.

Table 1: The symmetry classes and selection rules for the internal vibrations of the isolated cyclopropane molecule (16,17).

<table>
<thead>
<tr>
<th>Symmetry class</th>
<th>No. of fundamental vibrations in classes</th>
<th>Raman</th>
<th>Infrared</th>
</tr>
</thead>
<tbody>
<tr>
<td>$A'_{1}$</td>
<td>3</td>
<td>active</td>
<td>inactive</td>
</tr>
<tr>
<td>$A'_{2}$</td>
<td>1</td>
<td>inactive</td>
<td>inactive</td>
</tr>
<tr>
<td>$E'$</td>
<td>4</td>
<td>active</td>
<td>active</td>
</tr>
<tr>
<td>$A''_{1}$</td>
<td>1</td>
<td>inactive</td>
<td>inactive</td>
</tr>
<tr>
<td>$A''_{2}$</td>
<td>2</td>
<td>inactive</td>
<td>active</td>
</tr>
<tr>
<td>$E''$</td>
<td>3</td>
<td>active</td>
<td>inactive</td>
</tr>
</tbody>
</table>
Fig. 3 Mn⁺⁺⁺⁺ zeolite
a. at 293K after taking at 573K
b. with C₃H₆ at 293K
c. C₃H₆ gas
Table 2 contains some of the infrared and Raman frequencies of cyclopropane gas, liquid and from the crystals together with their assignments as given by Linnett (16).

**Table 2: Infrared and Raman data (cm\(^{-1}\)) of cyclopropane gas, liquid and crystals in the range of 3000-1000 cm\(^{-1}\) (16,17).**

<table>
<thead>
<tr>
<th>Infrared</th>
<th>Raman</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>-</td>
<td>2952  w</td>
<td>(4\nu_{14})</td>
</tr>
<tr>
<td>-</td>
<td>2854  w</td>
<td>(2\nu_9 (A_1^+E'))</td>
</tr>
<tr>
<td>2631 w</td>
<td>-</td>
<td>(\nu_3^+\nu_9 (E'))</td>
</tr>
<tr>
<td>2493 w</td>
<td>-</td>
<td>(\nu_2^+\nu_{10} (E'))</td>
</tr>
<tr>
<td>2330 w</td>
<td>-</td>
<td>(\nu_5^+\nu_9 (E'))</td>
</tr>
<tr>
<td>2178 w</td>
<td>-</td>
<td>(\nu_2^+\nu_{11} (E'))</td>
</tr>
<tr>
<td>2084 s</td>
<td>-</td>
<td>(\nu_9^+\nu_{14} (A_2'))</td>
</tr>
<tr>
<td>1888 s</td>
<td>1873 vW</td>
<td>(\nu_{10}^+\nu_{11} (A_1^+E'))</td>
</tr>
<tr>
<td>1779 w</td>
<td>-</td>
<td>(\nu_7^+\nu_{14} (E'))</td>
</tr>
<tr>
<td>1739 m</td>
<td>-</td>
<td>(\nu_4^+\nu_{14} (E'))</td>
</tr>
<tr>
<td>-</td>
<td>1504 w</td>
<td>(\nu_2 \text{CH}_2 (a_1'))</td>
</tr>
<tr>
<td>-</td>
<td>1454 m</td>
<td>(2 \nu_{14} (A_1^+E'))</td>
</tr>
<tr>
<td>1432 s</td>
<td>1435 m</td>
<td>(\nu_9 \text{CH}_2 (e))</td>
</tr>
<tr>
<td>-</td>
<td>1189 v.s</td>
<td>(\nu_3 \text{C}_3 (a_1'))</td>
</tr>
<tr>
<td>-</td>
<td>1120</td>
<td>(\nu_{13} (e^-)) rocking</td>
</tr>
<tr>
<td>1027.6 s</td>
<td>1022 w</td>
<td>(\nu_{10} (e)) bending</td>
</tr>
</tbody>
</table>
After adsorbing cyclopropane gas on each of the dehydrated Co-4A and Mn-4A zeolites at room temperature and evacuating for 10 minutes (Fig.5), new bands were observed in the 1500-1300 cm\(^{-1}\) region (Table 3). Two bands were observed, at 1456 and 1430 cm\(^{-1}\), for Co-4A zeolite and one band at 1435 cm\(^{-1}\) for Mn-4A zeolite.

Table 3: Infrared data and assignments of Mn-4A and Co-4A zeolites and the adsorbed species \((C_3H_6)\) in the range 1300-1700 cm\(^{-1}\)

<table>
<thead>
<tr>
<th>Mn-4A</th>
<th>Mn-4A.C(_3)H(_6)</th>
<th>Co-4A</th>
<th>Co-4A.C(_3)H(_6)</th>
<th>Assignment</th>
</tr>
</thead>
<tbody>
<tr>
<td>1660 s</td>
<td>1650 s</td>
<td>1650 s</td>
<td>1645 s</td>
<td>(v_2(H_2O))</td>
</tr>
<tr>
<td>-</td>
<td>-</td>
<td>1610 w</td>
<td>1610 w</td>
<td>zeolite structure</td>
</tr>
<tr>
<td>1535 s</td>
<td>1535 s</td>
<td>-</td>
<td>-</td>
<td>zeolite structure</td>
</tr>
<tr>
<td></td>
<td></td>
<td>1465 w</td>
<td>1465 w</td>
<td>zeolite structure</td>
</tr>
<tr>
<td>-</td>
<td>-</td>
<td>-</td>
<td>1456 m</td>
<td>(2v_{14}(A'_{1}+E'))</td>
</tr>
<tr>
<td>-</td>
<td>1435 m</td>
<td>-</td>
<td>1430 w</td>
<td>(v_9 (E))</td>
</tr>
<tr>
<td>1395 s</td>
<td>1395 s</td>
<td>1390 w</td>
<td>1390 w</td>
<td>zeolite structure</td>
</tr>
</tbody>
</table>

The band at 1456 cm\(^{-1}\) can be assigned to \(2v_{14}(A'_{1}+E')\) which is not observed in the infrared spectra with isolated cyclopropane molecules as can be seen from Table 2. This transition is observed because when cyclopropane is adsorbed on the zeolite the symmetry of the molecule is lowered from \(D_{3h}\) to \(C_{3v}\) (Table 4). Thus \(A_{1}(D_{3h})\) correlates with \(A_{1}(C_{3v})\) which is active in the infrared (Table 5).
Fig. 5

a. Mn-4A-\( \text{C}_3\text{H}_6 \) at 293K
b. Co-4A-\( \text{C}_3\text{H}_6 \) at 293K
c. \( \text{C}_3\text{H}_6 \) gas
Table 4: The symmetry classes and selection rules for adsorbed $\text{C}_3\text{H}_6$ molecules ($\text{C}_{3\nu}$)

<table>
<thead>
<tr>
<th>Symmetry class</th>
<th>No. of fundamental vibrations in classes</th>
<th>Raman</th>
<th>Infrared</th>
</tr>
</thead>
<tbody>
<tr>
<td>$A_1$</td>
<td>5</td>
<td>active</td>
<td>active</td>
</tr>
<tr>
<td>$A_2$</td>
<td>3</td>
<td>inactive</td>
<td>inactive</td>
</tr>
<tr>
<td>$E$</td>
<td>8</td>
<td>active</td>
<td>active</td>
</tr>
</tbody>
</table>

Table 5: Resolution of $D_{3h}$ into $C_{3\nu}$ (17)

<table>
<thead>
<tr>
<th>$D_{3h}$</th>
<th>$C_{3\nu}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>$A'_1$</td>
<td>$A_1$</td>
</tr>
<tr>
<td>$A'_1$</td>
<td>$A_2$</td>
</tr>
<tr>
<td>$A'_2$</td>
<td>$A_1$</td>
</tr>
<tr>
<td>$E'$</td>
<td>$E$</td>
</tr>
<tr>
<td>$E'$</td>
<td>$E$</td>
</tr>
</tbody>
</table>
The 1456 cm\(^{-1}\) band has not been observed with Mn-4A zeolite and this is because of the complexity of the spectra of Mn-4A zeolite compared with Co-4A zeolite in the 1600-1300 cm\(^{-1}\) region. The band at 1430 cm\(^{-1}\) which was observed in the spectra of Co-4A zeolite and 1435 cm\(^{-1}\) in the spectra of Mn-4A zeolite may be assigned to \(v_g(E)\) of the adsorbed C\(_3\)H\(_6\).

The interaction between cyclopropane and the zeolites can be described as Miller (18) discussed it using Hallmann-Feynman arguments and ring bending strain energies and they calculated a 'negative pole' in cyclopropane to be 2.88 e\(^-\). This negative pole or \(\pi\) cloud would be polarized by the field of the Co(II) or Mn(II) ions to form the bond between the cyclopropane and the transition metal ions. It has been found that cyclopropane is adsorbed by partially Ni(II)-exchanged zeolite A complexes (19). Ni(II) was found to be bonded to the centre of the ring, which involve \(\pi\) interaction.

4. **Conclusion**

From chapters 5 and 6 we can see that there are considerable differences in behaviour between partially exchanged Co-4A and Mn-4A zeolites. These differences include:

1. Mn-4A zeolite can form a hydronium ion when water vapour is adsorbed on the hot zeolite, while Co-4A zeolite does not.

2. Various spectroscopic studies (20) on adsorbed ethylene and cyclopropane on each of Mn-4A and Co-4A zeolites using neutron scattering have also shown the difference in behaviour between the zeolites.
These show that the Mn complexes are stronger than the Co complexes. Since Mn and Co are close together in the periodical table, one would expect their interactions with ligands to be fairly similar but in fact it is now known that this is not the case.

These differences between the Mn and Co zeolites once again illustrate the way in which zeolite properties can be influenced by subtle changes in the zeolite composition.
References


In chapter 4, a study of the relation between coke formation and hydroxyl groups which act as active sites was reported. This work was an attempt to repeat some work published earlier but which was unsuccessful because in our samples the band due to the relevant hydroxyl groups were not present at high temperature. Coke formation was, however, observed. Various studies on coke formation using different zeolites could be done to clarify whether or not there is an effect of any particular hydroxyl groups on coke formation.

The exchange of D$_2$O for H$_2$O by stirring the liquid with zeolite A has been shown to be very difficult. A comparison study could be achieved by exchanging D$_2$O for H$_2$O on a different type of zeolite to find a suitable way of exchanging D$_2$O inside zeolite A.

Mn-4A zeolite gives different infrared bands when ion exchanged to different degrees and this makes it a complicated and interesting material. The same thing could be done with higher degrees of Mn-exchange as well as on Co-4A zeolite since they show a different behaviour in many respects.

Zeolites can form novel chemical species since stable organometallic complexes of C$_3$H$_6$ or N$_3$H$_3$ have not been observed outside a zeolite cage. In fact does not exist in the pure form. These studies are worth doing on different zeolites to observe the conditions under which these species could be formed.
Also of importance is the study of hydronium ions which, as we have seen, forms with Mn-4A zeolite but does not form with Co-4A zeolite. The degree of the Mn-exchange also effects the presence of the hydronium ion. A similar study could be done on a variety of cations and different zeolites to try to understand the conditions under which the hydronium ion is formed.